



Part A. Use the criss-cross method to write the formulas produced from the listed ions.

	Cl^-	CO_3^{2-}	OH^-	SO_4^{2-}	PO_4^{3-}	NO_3^-
Na^+						
NH_4^+						
K^+						
Ca^{2+}						
Zn^{2+}						
Fe^{3+}						
Al^{3+}						
Co^{3+}						
Fe^{2+}						
Mg^{2+}						
H^+						

Part B. Write the names of the compounds formed in the above table.

	Cl^-	CO_3^{2-}	OH^-	SO_4^{2-}	PO_4^{3-}	NO_3^-
Na^+						
NH_4^+						
K^+						
Ca^{2+}						
Zn^{2+}						
Fe^{3+}						
Al^{3+}						
Co^{3+}						
Fe^{2+}						
Mg^{2+}						
H^+						



I. Name the following Ionic Compounds:

1. LiBr
2. $\text{CuC}_2\text{H}_3\text{O}_2$
3. PbSO_3
4. NaClO_3
5. CaC_2O_4
6. NaHSO_4
7. Hg_2Cl_2
8. NaHCO_3
9. NiBr_3
10. AuCl_3
11. KMnO_4

II. Name the following Covalent Compounds:

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|----------------------------|----------------------------|
| 1. CO_2 | 11. PCl_3 |
| 2. CO | 12. PCl_5 |
| 3. SO_2 | 13. NH_3 |
| 4. SO_3 | 14. SCl_6 |
| 5. N_2O | 15. P_2O_5 |
| 6. NO | 16. CCl_4 |
| 7. N_2O_3 | 17. SiO_2 |
| 8. NO_2 | 18. CS_2 |
| 9. N_2O_4 | 19. OF_2 |
| 10. N_2O_5 | 20. PBr_3 |