

Properties and Trends – Supplemental Worksheet

1. What 2 prominent trends regarding ionization energy do we find from the periodic table?

2. Explain how shielding varies between outer electrons and core electrons and between outer electrons and other outer electrons.

3. Explain electron affinity.

4. What atomic radii trends are found on the periodic table? Explain why.

5. Using the element Nitrogen, write equations that correspond with electron affinity and ionization energy.

6. Arrange the following groups of atoms in order of decreasing size.

a. P, Sb, N, As b. Sr, In, Mo, Y c. F, Fr, Os, Ga d. Na, W, Si, Pb

7. Which atom or ion has the largest ionization energy?

a. Mg, Sr, Ba b. Ca, Co, Se c. O²⁻, O, O²⁺