

1. What is the percent by mass of a solution made by dissolving 5.25 g of calcium nitrate in 675 g of water?
2. How many grams of NaCl are present in 1250. g of solution that is 12.00% NaCl by mass?
3. What mass of water is contained in 600. g of 12.0 % NaCl solution?
4. What mass of water must we add to 35.0 g NaCl to make a 12.0 % NaCl solution?
5. What mass of 12.0 % NaCl solution contains 35.0 g of NaCl?
6. What mass of 25 % calcium chloride solution contains 350. g of water?
7. What mass of each calcium chloride and water are required to prepare 350. g of 22.0 % calcium chloride solution?
8. A solution is 12 % calcium hydroxide. How many moles of calcium hydroxide are dissolved in 250 g of this solution?
9. The density of a 15.00 % NaCl solution is 1.108 g/mL. How many mL of this solution must we use to obtain 75.00 g NaCl?
10. What is the molarity of a 20.00 % solution of NaNO_3 ? The density of the solution is 1.143 g/mL.
11. What is the density of a 35.0 % hydrochloric acid solution, HCl, if it's 11.3 molar?
12. The density of a 16.4 M NaOH solution is 1.43 g/mL. What is the percent by mass of this solution?