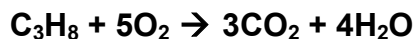


Consider the following chemical equation to answer the questions below.



- _____ molecules of C_3H_8 react with _____ molecules of O_2 to produce _____ molecules of CO_2 and _____ molecules of H_2O .
- 20 molecules of C_3H_8 react with _____ molecules of O_2 to produce _____ molecules of CO_2 and _____ molecules of H_2O .
- 6.022×10^{23} molecules of C_3H_8 react with _____ molecules of O_2 to produce _____ molecules of CO_2 and _____ molecules of H_2O .
- 1 mole of C_3H_8 reacts with _____ moles of O_2 to produce _____ moles of CO_2 and _____ moles of H_2O .
- 5 moles of C_3H_8 reacts with _____ moles of O_2 to produce _____ moles of CO_2 and _____ moles of H_2O .
- 44 g of C_3H_8 reacts with _____ grams of O_2 to produce _____ grams of CO_2 and _____ grams of H_2O .
- 176 g of C_3H_8 reacts with _____ grams of O_2 to produce _____ grams of CO_2 and _____ grams of H_2O .
- How many moles of carbon dioxide can be produced from the reaction of 12 moles of propane?
- What mass of oxygen is needed to produce 65 grams of water?
- How many moles of carbon dioxide can be produced from the reaction of 225 g of propane?
- What mass of H_2O is produced from the reaction of 6.3 g of propane?
- How many molecules of H_2O are produced when 2 moles of O_2 are reacted with excess propane?