

Standard Potentials at 25°C

(v4 : 4-29-13)

Half Reaction	Potential	Half Reaction	Potential
$F_2 + 2e^- \rightleftharpoons 2F^-$	+2.87 V	$2H^+ + 2e^- \rightleftharpoons H_2$	0.000 V
$O_3 + 2H^+ + 2e^- \rightleftharpoons O_2 + H_2O$	+2.07 V	$Fe^{3+} + 3e^- \rightleftharpoons Fe$	-0.04 V
$S_2O_8^{2-} + 2e^- \rightleftharpoons 2SO_4^{2-}$	+2.05 V	$Pb^{2+} + 2e^- \rightleftharpoons Pb$	-0.13 V
$H_2O_2 + 2H^+ + 2e^- \rightleftharpoons 2H_2O$	+1.78 V	$Sn^{2+} + 2e^- \rightleftharpoons Sn$	-0.14 V
$PbO_2 + 4H^+ + SO_4^{2-} + 2e^- \rightleftharpoons PbSO_4 + 2H_2O$	+1.69 V	$Ni^{2+} + 2e^- \rightleftharpoons Ni$	-0.23 V
$Au^+ + e^- \rightleftharpoons Au$	+1.69 V	$V^{3+} + e^- \rightleftharpoons V^{2+}$	-0.26 V
$Pb^{4+} + 2e^- \rightleftharpoons Pb^{2+}$	+1.67 V	$Co^{2+} + 2e^- \rightleftharpoons Co$	-0.28 V
$2 HClO + 2H^+ + 2e^- \rightleftharpoons Cl_2 + 2H_2O$	+1.63 V	$In^{3+} + 3e^- \rightleftharpoons In$	-0.34 V
$Ce^{4+} + e^- \rightleftharpoons Ce^{3+}$	+1.61 V	$PbSO_4 + 2e^- \rightleftharpoons Pb + SO_4^{2-}$	-0.36 V
$MnO_4^- + 8H^+ + 5e^- \rightleftharpoons Mn^{2+} + 4H_2O$	+1.51 V	$Cd^{2+} + 2e^- \rightleftharpoons Cd$	-0.40 V
$Au^{3+} + 3e^- \rightleftharpoons Au$	+1.40 V	$Cr^{3+} + e^- \rightleftharpoons Cr^{2+}$	-0.41 V
$Cl_2 + 2e^- \rightleftharpoons 2Cl^-$	+1.36 V	$Fe^{2+} + 2e^- \rightleftharpoons Fe$	-0.44 V
$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightleftharpoons 2Cr^{3+} + 7H_2O$	+1.33 V	$U^{4+} + e^- \rightleftharpoons U^{3+}$	-0.61 V
$O_2 + 4H^+ + 4e^- \rightleftharpoons 2H_2O$	+1.23 V	$FeCO_3 + 2e^- \rightleftharpoons Fe + CO_3^{2-}$	-0.756 V
$MnO_2 + 4H^+ + 2e^- \rightleftharpoons Mn^{2+} + 2H_2O$	+1.21 V	$Zn^{2+} + 2e^- \rightleftharpoons Zn$	-0.76 V
$Pt^{2+} + 2e^- \rightleftharpoons Pt$	+1.20 V	$2H_2O + 2e^- \rightleftharpoons H_2 + 2OH^-$	-0.83 V
$Br_2 + 2e^- \rightleftharpoons 2Br^-$	+1.09 V	$Cr^{2+} + 2e^- \rightleftharpoons Cr$	-0.91 V
$Pd^{2+} + 2e^- \rightleftharpoons Pd$	+0.915 V	$Mn^{2+} + 2e^- \rightleftharpoons Mn$	-1.18 V
$2Hg^{2+} + 2e^- \rightleftharpoons Hg_2^{2+}$	+0.92 V	$V^{2+} + 2e^- \rightleftharpoons V$	-1.19 V
$ClO^- + H_2O + 2e^- \rightleftharpoons Cl^- + 2OH^-$	+0.89 V	$ZnS + 2e^- \rightleftharpoons Zn + S^{2-}$	-1.44 V
$Ag^+ + e^- \rightleftharpoons Ag$	+0.80 V	$Al^{3+} + 3e^- \rightleftharpoons Al$	-1.66 V
$Hg_2^{2+} + 2e^- \rightleftharpoons 2Hg$	+0.79 V	$Mg^{2+} + 2e^- \rightleftharpoons Mg$	-2.36 V
$Fe^{3+} + e^- \rightleftharpoons Fe^{2+}$	+0.77 V	$Na^+ + e^- \rightleftharpoons Na$	-2.71 V
$MnO_4^- + 2H_2O + 3e^- \rightleftharpoons MnO_2 + 4OH^-$	+0.60 V	$K^+ + e^- \rightleftharpoons K$	-2.92 V
$I_2 + 2e^- \rightleftharpoons 2I^-$	+0.54 V	$Li^+ + e^- \rightleftharpoons Li$	-3.05 V
$O_2 + 2H_2O + 4e^- \rightleftharpoons 4OH^-$	+0.40 V		
$Cu^{2+} + 2e^- \rightleftharpoons Cu$	+0.34 V		
$Hg_2Cl_2 + 2e^- \rightleftharpoons 2Hg + 2Cl^-$	+0.27 V		
$AgCl + e^- \rightleftharpoons Ag + Cl^-$	+0.22 V		
$Bi^{3+} + e^- \rightleftharpoons Bi$	+0.20 V		
$NO_3^- + H_2O + 2e^- \rightleftharpoons NO_2^- + 2OH^-$	+0.01 V		
$2H^+ + 2e^- \rightleftharpoons H_2$	0.000 V		

Note: all ions are aqueous (aq), many neutral species are solids (s), although some are liquids (l), gases (g), and even aqueous (aq). Use other sources for details on state. They were purposely left off here to save space and keep a cleaner looking table.