

## Common Polyatomic Ions

Name(s)	Formula	Name(s)	Formula
ammonium	$\text{NH}_4^+$	iodate	$\text{IO}_3^-$
acetate	$\text{CH}_3\text{COO}^-$ $\text{C}_2\text{H}_3\text{O}_2^-$	nitrate	$\text{NO}_3^-$
bromate	$\text{BrO}_3^-$	nitrite	$\text{NO}_2^-$
carbonate	$\text{CO}_3^{2-}$	oxalate	$\text{C}_2\text{O}_4^{2-}$
chlorate	$\text{ClO}_3^-$	perchlorate	$\text{ClO}_4^-$
chlorite	$\text{ClO}_2^-$	periodate	$\text{IO}_4^-$
chromate	$\text{CrO}_4^{2-}$	permanganate	$\text{MnO}_4^-$
cyanide	$\text{CN}^-$	peroxide	$\text{O}_2^{2-}$
dichromate	$\text{Cr}_2\text{O}_7^{2-}$	phosphate	$\text{PO}_4^{3-}$
hydrogen carbonate bicarbonate	$\text{HCO}_3^-$	phosphite	$\text{PO}_3^{3-}$
hydrogen sulfate bisulfate	$\text{HSO}_4^-$	silicate	$\text{SiO}_4^{4-}$
hydrogen phosphate biphosphate	$\text{HPO}_4^{2-}$	sulfate	$\text{SO}_4^{2-}$
hydroxide	$\text{OH}^-$	sulfite	$\text{SO}_3^{2-}$
hypochlorite	$\text{ClO}^-$	thiocyanate	$\text{SCN}^-$
		thiosulfate	$\text{S}_2\text{O}_3^{2-}$

## Other Ions

copper (I)	cuprous	$\text{Cu}^+$
copper (II)	cupric	$\text{Cu}^{2+}$
iron (II)	ferrous	$\text{Fe}^{2+}$
iron (III)	ferric	$\text{Fe}^{3+}$
lead (II)	plumbous	$\text{Pb}^{2+}$
lead (IV)	plumbic	$\text{Pb}^{4+}$
mercury (I)	mercurous	$\text{Hg}_2^{2+}$
mercury (II)	mercuric	$\text{Hg}^{2+}$
tin (II)	stannous	$\text{Sn}^{2+}$
tin (IV)	stannic	$\text{Sn}^{4+}$